## TREEZE WAIER F

## **Materials**

- 1 Glass thermometer
- 1 Metal-backed thermometer
- 1 Plastic cup
- 1 Vial

- 1 Stirring stick
- Sodium chloride, 3 spoons
- Ice, crushed
- Protective eyewear



Water

## **Procedure**

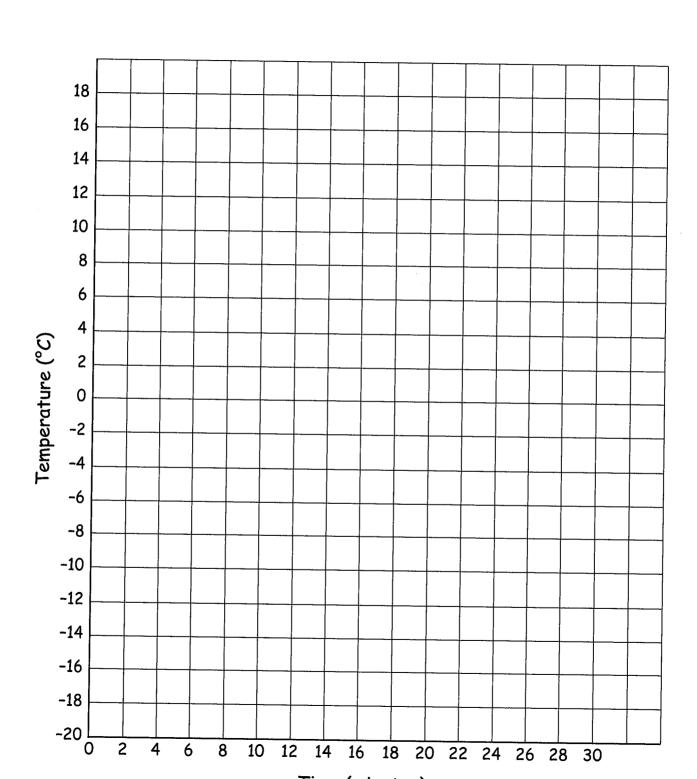
- a. Fill a plastic cup halfway with crushed ice.
- b. Put on protective eyewear. Add three 5-mL spoons of sodium chloride to the ice. Stir in thoroughly.
- c. Put about 10 mL of water in a vial.
- d. Carefully work the vial of water into the crushed ice.

  Make sure the surface of the water is below the level of the ice.
- e. Monitor the temperature of the water in the vial with a glass thermometer.

  Monitor the temperature of the ice/salt environment with a metal-backed thermometer.
- f. Record your observations. Include time, temperatures, and changes to the system.

## **Results**

Time	Water temperature (°C)	Ice bath temperature (°C)	Observations



	Describe what happened to the ice/salt mixture as the investigation progressed. <i>Explain using proper vocabulary</i> .			
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	What happened to the temperature of the water in the vial as the water we freezing? <u>Make sure to explain what is happening at the molecular scale Al use proper vocabulary</u> .			
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4. Based on your observations in this activity, why do people put salt on ice when

they make ice cream?

Conclusion questions- All answers must be in complete sentences. Remember, energy